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Answer | Std 9 |

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**BASES \u0026amp;**

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**CHAPTER || CLASS**

**10 CBSE**

**CHEMISTRY Acid**

**Bases \u0026amp;**

**salts chapter 5**

**QUESTION ANSWERS**

**full**

**explanation**

**class 7th ncert**

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Solutions (Part  
2) - Acid, Bases  
And Salts |

Class 10  
Chemistry

9th std Science  
5. Acids, Bases  
and Salts

OBJECTIVE

QUESTIONS \u0026amp;

ANSWERS Class 7th

Science Chapter

5: Acids, Bases

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Question-Answers  
(English Medium)

*Acids, Bases And  
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CBSE Class 10*

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**solutions for  
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| acids bases  
and salts class  
10 exercise  
questions**



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**Answers** *Acids,  
Bases and Salts  
| Ch-2, Part-3 |  
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*explained in  
hindi **Acid ,***

**Base and Salt ??**

??? 39 ?????? ||

**Lucent Objective**

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*Acids, Bases*

*\u0026 Salts - 1*

*| Class 10*

*Chemistry |*

*Science Chapter*

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Questions pH ???

????? | pH value  
of important

Substance | Gk  
for railway ALP,  
Group D 2018

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~~Acids bases and  
salts complete  
exercise solutio  
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bases and salts-  
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Acid, Base, Salt  
| 50 Important  
MCQ ~~Acids Bases~~*

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bases and salts  
- Part 2 |

Explanation in  
Hindi **Questions  
and answers on  
acid, base, and  
salt** || ????,  
??????, ???

Q\u0026A ||  
**Class 10 SCP**

Acids, Bases and

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Salts | Class 7

Science Sprint  
for Final Exams

| Chapter 5

@Vedantu Young

Wonders ~~Acids~~

~~Bases And Salts~~

Answers

Solution: The  
reaction of the  
acid + base  
gives a product  
of salt + water,  
which is

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Bases And as  
neutralization  
Salts Answers  
reaction.

Examples:  $\text{NaOH} +$   
 $\text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

~~NCERT Solutions  
for class 10  
chapter 2 Acids,  
bases and salt~~  
Class 10 Science  
Chapter 2 Extra  
Questions and



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Bases and Salts.

Extra Questions

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2 Acids Bases

and Salts with

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Solutions. Acids

Bases and Salts

Extra Questions

Very Short

Answer Type.

Question 1. Two

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~~Solutions And~~  
~~Salts Answers~~  
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solutions A and  
B have pH values  
of 5 and 8  
respectively.

~~Chemistry~~  
Which solution  
will be basic in  
nature ...

~~Acids Bases and~~  
~~Salts Class 10~~  
~~Extra Questions~~  
~~with ...~~

Ans:- Salts are  
classified into

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Chemistry

four categories  
on the basis of  
the strength of  
acid and base.

they are: Salt  
of strong acid  
and strong  
base. e.g NaCl,  
KCl; Salt of  
strong acid and  
weak base. e.g.  
NH<sub>4</sub> Cl, NH<sub>4</sub> I;  
Salt of weak  
acid and strong

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base. e.g.  $\text{CH}_3\text{COONa}$ ,  $\text{COO}^- \text{Na}^+$  + Salt of weak acid and weak

base. e.g.  $\text{CH}_3\text{COO}^- \text{NH}_4^+$  + How is common salt found in nature?

~~Class 10 Science  
Acids, Bases and  
Salts — SEBA +  
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Bases And exam  
revision with  
multiple choice  
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for Acids, Bases  
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& model answers

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Bases Acids, Bases  
& Oxides. . . .

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Salts. 8.1

Chemistry  
Acids, Bases &  
Oxides; 8.2

Salts & Chemical  
Analysis; 9. The  
Periodic Table

. . .

~~Acids, Bases &  
Oxides | CIE  
IGCSE Chemistry~~

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~~1. MCQ & Answers~~

2. Bases react with acids to produce salts and water; water to produce acids and salts; salts to produce acids and water; neither acids, salts, nor water

~~Acids Bases and  
Salts Multiple~~

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~~Choice Questions  
and Answers~~

The salts of strong acids and weak bases give acidic solution having pH less than 7. Example,  $\text{NH}_4\text{Cl}$ , Ammonium Chloride will have pH less than 7. The salts of weak



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Acids and strong  
bases give basic  
solution having  
pH more than 7.

Example,  $\text{Na}_2\text{CO}_3$ , Sodium  
Carbonate will  
have pH more  
than 7. Question  
7.

~~Acids, Bases and  
Salts Class 10  
Important~~

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Answer Key. Some

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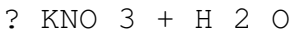
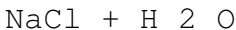
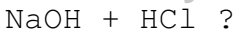
Answer A  
reaction in  
which an acid  
and base react  
with each other  
to give a salt

# Where To Download Acids

and water is

termed as  
neutralization  
reaction. For

Example: (i)



15. Give two  
important uses  
of washing soda  
and baking soda.

Answer Two

# Where To Download Acids

important uses  
of washing soda  
are:

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Science: Ch 2  
Acids, Bases ...~~

Answer. (a) HCl  
is acid and NaOH  
is base whose  
combination  
forms the common  
salt. Its

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Formula is  $\text{NaCl}$   
(Sodium  
chloride). It is  
obtained from  
sea water. (b)  
Rock salt is the  
common name for  
the mineral  
"halite". Its  
chemical formula  
is  $\text{NaCl}$ . It may  
be white or  
light blue or  
yellow depending

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Bases and  
Impurities  
present in it.

~~Acids Bases and~~

~~Salts Chapter~~

~~Wise Important~~

~~Questions ...~~

Acids Bases And

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Taking Answers

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Inhibitors  
Corrosion Rust.  
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Exam Pattern.

phenylamine  
aniline as an  
amine chemguide.

## Chemistry

~~Acids Bases And  
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Taking Answers~~

(A) inorganic  
acid (B) organic  
acid (C) soft  
organic acid (D)  
strong inorganic  
acid. Answer.

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Answer: (a) 5.

When milk of magnesia reacts with acetic acid it produces (A) basic salts (B) acidic salts (C) neutral salt (D) complex salt.

Answer. Answer: (c) 6. Which of the following phenomena will occur when a

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Small amount of  
acid is added to  
water ...

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and Answers~~

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Salts Answer Key

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Multiple Choice

Questions with

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Choice Questions  
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know their  
preparation  
level.

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Bases and Salts~~

...

Turmeric is a  
natural  
indicator. On  
adding its paste  
to acid and base



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separately,

which colours  
would be

observed. (a)

Yellow in both  
acid and base

(b) Yellow in  
acid and red in  
base (c) Pink in  
acid and yellow  
in base (d) Red  
in acid and blue  
in base. Answer.

Answer: (b)

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Bases in acid  
and red in base

# Salts Answers

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Answers. Choose  
the correct  
answer: Question

1.

## Chemistry

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Acids, Bases and

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first quiz on

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we take a look  
at one vital  
part of the  
subject – acids,  
alkalis, bases  
and salts. In it  
we find out the  
difference  
between alkalis  
and bases, how  
acids and  
alkalis react  
with one  
another, how

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neutralisation  
can produce  
salts and the  
relevance of pH  
values.

~~GCSE Acids,  
Bases and Salts  
+ Revise pH  
Values and  
Alkalis~~

Acids, bases and  
salts An acid is  
a compound that

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is defined by  
its physical and  
chemical  
properties.

Acids taste sour  
and react with  
metals and  
polyatomic ions  
called  
carbonates. A  
carbonate is a  
charged cluster  
of Carbon and  
Oxygen atoms.

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salts. 7th Grade  
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well.

## Salts Answers

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~~2 Acids, Bases~~

~~...~~

Acid + Base ---->  
Salt + Water.

Question 3: Rain  
becomes acidic  
due to. carbon  
dioxide. sulphur

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dioxide. And  
nitrogen  
dioxide. all of  
the above.

Correct Option  
is : 4. Solution  
: Rain becomes  
acidic because  
of carbon  
dioxide, sulphur  
dioxide and  
nitrogen  
dioxide. They  
dissolve in rain

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drops to form  
carbonic acid,  
sulphuric acid  
and nitric acid

•••  
Chemistry

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