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Answers CHAPTER 11 SOLUTIONS
MANUAL Section 11.1 Defining
Stoichiometry pages 368–372
Practice Problems pages
371–372 1. Interpret the
following balanced chemical
equations in terms of
particles, moles, and mass.

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Show that the law of
conservation of mass is
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...

Stoichiometry Chapter 11

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In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and

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the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

Chapter 11.4: Stoichiometry

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15.2 CHAPTER 11:

STOICHIOMETRY. MOLE TO MOLE
RATIO. When nitrogen and
hydrogen gas are heated
under the correct
conditions, ammonia gas
(NH_3) is formed. a. RXN: 1.

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$N_2 + 3 H_2 \rightarrow 2 NH_3$. b. How many moles of nitrogen react with three moles of

hydrogen? 1 mol N_2 3 mol H_2 1 mol N_2 . 3 mol H_2 .

c. How many moles of nitrogen react with six moles of hydrogen? 2 mol

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N_2 _____ 6 mol H_2 1 mol N_2

...

CHAPTER 11: STOICHIOMETRY

Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts

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of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical

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Stoichiometry is the tool
for answering these
questions. Stoichiometry The
study of quantitative

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relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass.

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stoichiometry. mole ratio.

excess reactant. limiting

reactant. The study of

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quantitative relationships between the amounts of... In a balanced equation, the ratio between the numbers of moles... A reactant that remains after a chemical reaction stops.

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The ratio of the amount of product that is expected to be produced in a chemical reaction (using stoichiometric calculations based upon the balanced

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chemical equation) and the amount of product actually produced when carried out in an actual experiment.

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Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining Stoichiometry Vocab • stoichiometry • mole ratio Concepts Using Balanced Equations • Number of Atoms

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