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Stoichiometry Exercises. Answer the following to the best of your ability. Questions left blank are not counted against you. When you have completed every question that you desire, click the "MARK TEST" button after the last exercise at the bottom of the page. A new page will appear showing your correct and incorrect responses.

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## Stoichiometry Exercises

Stoichiometry Chapter 3! Stoichiometry: Calculations with Chemical Formulas and Equations. Stoichiometry Anatomy of a Chemical Equation  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  Stoichiometry Anatomy of a Chemical Equation Reactants appear on the left side of the equation.  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  ...

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Chapter 3 Stoichiometry - Home - Chemistry  
CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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mc06se cFMsr i-vi

Exercise-1: 1. Naturally occurring magnesium is composed of 78.99% of  $^{24}\text{Mg}$  (23.9850 u), 10.00%  $^{25}\text{Mg}$  (24.9858 u), and 11.01%  $^{26}\text{Mg}$  (25.9826 u). Calculate the average atomic mass of magnesium. (Answer: 24.31 u) 2. Chlorine has two naturally occurring isotopes, one with a mass of 34.96885 u and the other a mass of 36.96590 u.

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## CHAPTER 3 - STOICHIOMETRY

Clark, Smith (CC-BY-4.0) GCC CHM 130 Chapter 13: Stoichiometry page 1 Chapter 13 - Stoichiometry Stoichiometry (STOY-key-OM-etry) problems are based on quantitative relationships between the ... 3 (g) Answer: 36.5 L  $\text{SO}_3$  (1 mol  $\text{SO}_3$  22.4 L  $\text{SO}_3$ ) (1 mol  $\text{O}_2$  2 mol  $\text{SO}_3$ ) (22.4 L  $\text{O}_2$  1 mol  $\text{O}_2$ ) = 18.3 L  $\text{O}_2$  (notice molar volume cancels out with itself ...

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Chapter 13 Stoichiometry - Welcome to [web.gccaz.edu](http://web.gccaz.edu)

The general approach to using stoichiometric relationships is similar in concept to the way people go about many common activities. Food preparation, for example, offers an appropriate comparison. A recipe for making eight pancakes calls for 1 cup pancake mix, 3/4 cup milk, and one egg. The "equation" representing the preparation of ...

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