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Configuration

for an Element

in Each Block

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Molecular

Compounds | How

to Pass

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#5 The Origin of

the Elements

Atomic Hook-Ups

- Types of

Chemical Bonds:

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Energy Levels,

Energy

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Orbitals, \u0026

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Principle

Partial

Pressures \u0026

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Gas Unit

Properties
of a Gas. Gas
Balloon. ... A
gas is a state

of matter. The
particles that
make up a gas
can range
from...

Pressure.

Pressure is a
measure of the
amount of force
per unit area.

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The pressure of
a gas is the
amount of force
the...

Temperature.

Temperature is a
property of ...

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The ideal gas

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law is expressed

by the formula

$PV = nRT$ where P

= pressure V =

volume n =

number of moles

of gas R = ideal

gas constant T =

absolute

temperature The

value of R

depends on the

units of

pressure, volume

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and temperature.

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The study of
gases allows us
to understand
the behavior of
matter at its

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simplest: Unit

individual
particles,
acting

independently,

almost

completely

uncomplicated by

interactions and

interferences

between each

other.

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Chemistry Gases

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A gas is a state of matter with no defined shape or volume. Gases have their own unique behavior depending on a variety of variables, such as temperature,

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pressure, and
volume. While
each gas is
different, all
gases act in a
similar matter.

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particles per
unit area What
is the formula
for gas
pressure? $P = F/A$

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where P =pressure
and F =force of
the gas
particles and
 A =area What are
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Unit 2: States
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1. Oxygen and
Hydrogen combine
to form water.

In the space to
the right, draw
an illustration

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from the Ideal

Gas Law In the

real world,

gases don't

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always behave as
defined by the
Kinetic
Molecular
Theory.

Conditions of
high pressure
and low
temperature will
cause gases to
deviate from
ideal gas
behavior for the
following

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gases, you can
convert mass to
moles (or moles
to mass). For
gases, you
should memorize
the Page 9/14

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