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// THEORIES IN CHEMICAL KINETICS // CHEMICAL KINETICS // Chemical Kinetics | CBSE | Class 12 Chemistry | NCERT | Collision Theory Gizmo Lab Activity A (L-18) Effect of Catalyst \u0026 Collision Theory of

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Collision Theory. Launch Gizmo. Observe a chemical reaction with and without a catalyst. Determine the effects of concentration, temperature, surface area, and catalysts on reaction rates. Reactant and product concentrations through time are recorded, and the speed of the simulation can be adjusted by the user. Launch Gizmo.

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DESCRIPTION Observe a chemical reaction with and without a catalyst. Determine the effects of concentration, temperature, surface area, and catalysts on reaction rates. Reactant and product concentrations through time are recorded, and the speed of the simulation can be adjusted by the user.

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Collision theory answer key exploration guide collision theory gizmo answer key. Reactant and product concentrations through time are recorded and the speed of the simulation can be adjusted by the user. I have noticed this year that many students skip these questions and it hurts their ability to connect to what the lesson is

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What is the collision theory? It is the idea that particles have to collide in order to react and they have to collide hard enough (with enough energy to break the bonds - activation energy). How does a higher temperature increase the rate of reaction? When the temperature is increased the particles have more energy and move quicker.

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HS-PS1-5: Collision Theory and Rates of Reaction Apply scientific principles and evidence to provide an explanation about the effects of changing the temperature or concentration of the reacting particles on the rate at which a reaction occurs .

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