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15. 0.150 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?  $2 \text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$  2. How many mL of 0.

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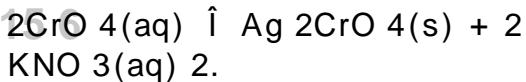
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0.150 L  $\text{AgNO}_3$  0.500

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moles AgNO<sub>3</sub> 1 moles Ag<sub>2</sub>CrO<sub>4</sub>  
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replacement reactions involve the  
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in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous solution, we can use the concept of molarity to directly calculate the number of



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