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first law as  $w = q - DU$  ;  
thus  $q = DH$ ;  $w = D$

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Process Because  $U$  is a  
function only of  $T$  for an  
ideal gas,  $DU = DH = 0$   
for an isothermal  
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 $DT = D(PV) = 0$  for an  
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of  $V = V(P, T)$  for a gas is shown in Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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the values of  $T$ ,  $V$  and  $P$   
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Step 1 T V P A T V P B

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first law as  $w = q - \Delta U$  ;

thus  $q = \Delta H$ ;  $w = \Delta D$

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