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A chemical bond is formed when electrons are shared between two atoms. There are three types of bonds: covalent bonds, polar covalent bonds and ionic bonds.

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The simplest example of bonding is demonstrated by the H_2 molecule.

We can see that each hydrogen atom has a single electron from the periodic table.

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7.2 IONIC BONDING

The chemical bond formed by transfer of electron from a metal to a non- metal is known as ionic or electrovalent bond . For example, when sodium metal and chlorine gas are brought into contact, they react

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violently and we obtain sodium chloride. This reaction is shown

below: $2\text{Na(s)} + \text{Cl}_2\text{(g)}$
 2NaCl(s)

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*7 CHEMICAL
BONDING*

There are many types of chemical bonds that can form, however the 3 main types are: ionic, covalent, and metallic bonds. You must

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Become familiar with how they work and the differences between the 3 types. I. Ionic bonding: Model 1 is a description of what chemists call ionic bonding. Ionic bonding occurs strictly between metal and nonmetal atoms.

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Accordingly, there are
different types of
chemical bonds, like,
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In addition to these we have a special kind of bond called hydrogen bond. Let us discuss about different types of bonds, their formation and the properties of the compounds so formed.

5.2 Ionic or Electrovalent Bond

CHEMICAL BONDING

There are two main types and some

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secondary types of chemical bonds: 1 Ionic bond Ionic bonding involves a transfer of an electron, so one atom gains an electron while one atom loses an electron. One of the resulting ions carries a negative charge (anion), and the other ion carries a positive charge (cation).

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